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 $N_2 + 3H_2 \rightarrow 2NH_3$  What volume of  $NH_3$   
at STP is produced if 25.0 of  $N_2$  is  
reacted with an excess of  $H_2$ ? 3 3 3 2 3  
2 2 2 40.0L  $NH_3$  1mol  $NH_3$  22.4L  $NH_3$  ...  
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Answers your answer. Example 2:  
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255.507 mL The answer is expressed as  
255.5 mL since 46.0 mL has only  
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an excess of  $\text{O}_2$ ?  
 $2\text{N}_2 + 3\text{O}_2 \rightarrow 2\text{NO}_2$   
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PROBLEMS I.  $\text{N}_2 + 3\text{H}_2$  Name How many  
moles of hydrogen are needed to  
completely react with two moles

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